



Chemistry

Class-9

Chapter-6

Concept of Mole and Chemical counting

Subject teacher- Syeeda Sultana

Revision Work sheet -1

Date-06.10.2020

Exercise:

Write down the answers of the following questions on your copy.

Questions:

1. What is mole?
2. What is Avogadro number?
3. Determine the molecular mass of Na_2SO_4 ?
4. What is the relative molecular mass of glucose?
5. What is the mass of 1 H_2O molecule?
6. What is the mass of 5 water molecules?
7. How many H_2SO_4 molecules are there in 1g H_2SO_4 ?
8. Calculate the mole of H_2O present in 5g H_2O .
9. Determine the number of H,S and O atoms in 1g H_2SO_4 .
10. What is the total number of atoms present in 1 gram of methane (CH_4)?
11. What is the mass of 3.01×10^{23} atoms of carbon?
12. What is molar volume?
13. What is STP?
14. How many molecules are there in 1liter CO_2 gas at standard temperature and pressure?
15. How many H atoms are there in 5 liter CH_4 gas at standard condition?
16. What is the volume of 5 moles CO_2 gas at STP?
17. Calculate the volume of 5 CO_2 molecules at STP.
18. What is the volume of 10g hydrogen gas at STP?
19. What is molar solution?
20. What is molarity?
21. What is meant by semimolar and decimolar solution?
22. Molarity depends on temperature-Explain.
23. Prepare 2L 0.3M NaCl solution.
24. To prepare 2L 0.5M H_2SO_4 solution, what amount of solute should be needed?
25. Prepare 2 liter 0.1M NaHCO_3 solution.
26. Find out the volume of the solution when 20g Na_2CO_3 is dissolved to make 0.75M solution.